

Metaphysics Simulation Story: Defending Automotive Components Against Corrosive Destruction

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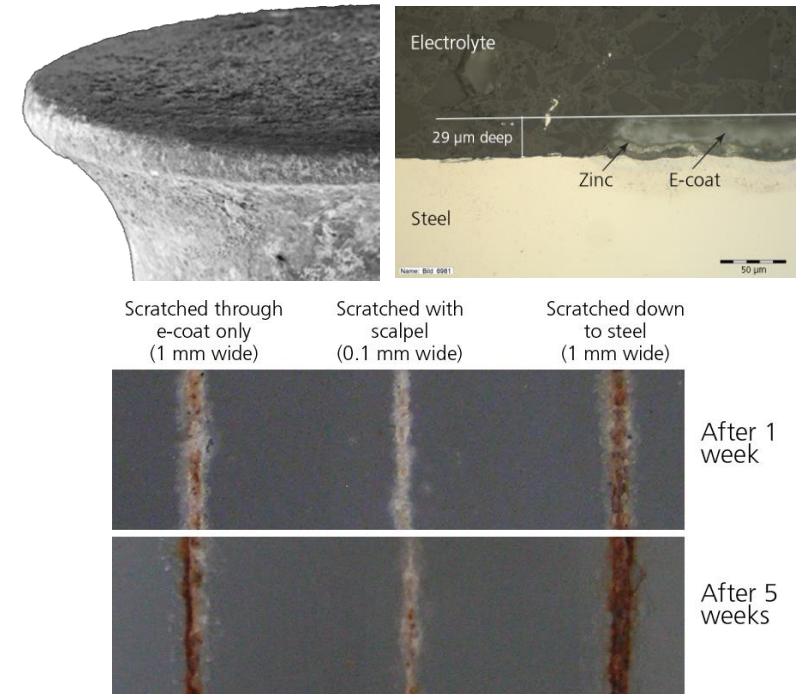
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The Challenge

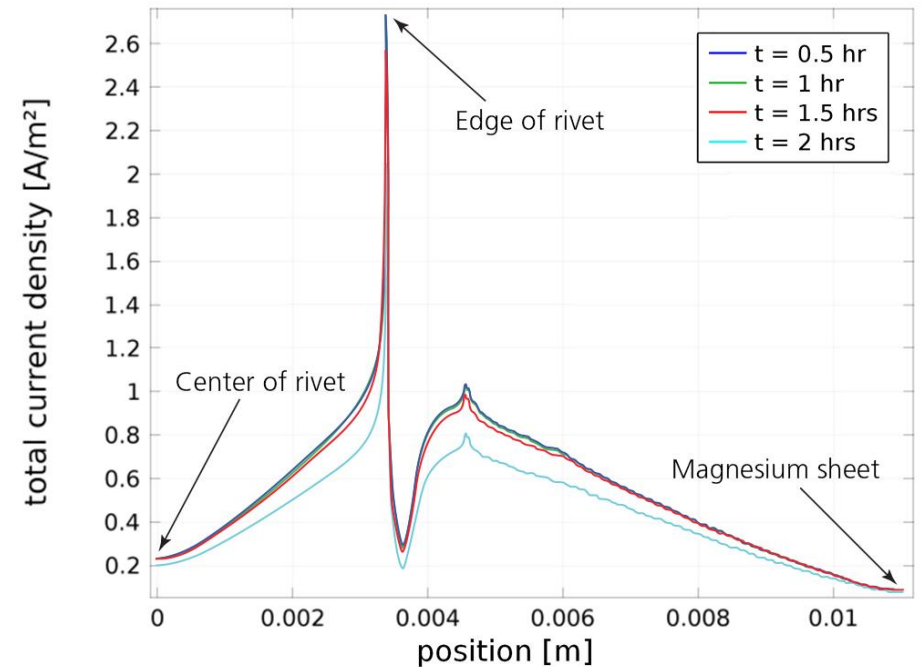
- Galvanic corrosion, caused by chemical reactions between different metals in contact with each other, weakens and eventually degrades areas of sheet metal and rivets in automotive components
 - Material properties, surface roughness, and joining techniques affect the chemical reactions
- Minor defects such as scratches can cause delamination of protective coatings, weakening corrosion protection
- Goal: redesign rivet geometry to minimize corrosion risk



Top left: rivet showing magnesium hydroxide deposit due to corrosion. Top right: cross-section of a test sheet where a scratch has destroyed part of the protective e-coat and zinc. Bottom: corrosion on a galvanized sheet over time.

The Solution

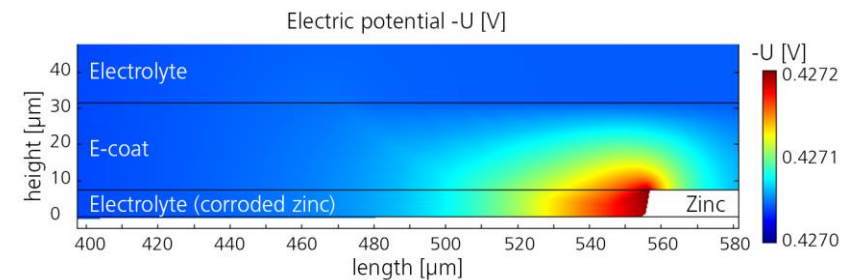
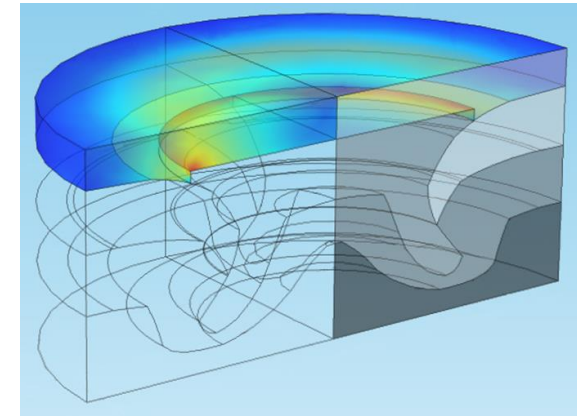
- Identify which geometric factors have the greatest influence on corrosion risk
 - Adjust geometry to minimize corrosion
 - Streamline rivet design and development
- Study variables affecting delamination of coatings that weaken protection
 - Investigate the impact of different sizes and depths of defects in sheet metal



COMSOL plot showing the localized current density at different positions on the surface of a rivet joint.

The Simulation

- Current flow analysis at the interface between rivet and sheet metal
 - Non-constant growth and material properties' influence
 - Predicted decay of adjoined sheets
 - Reduced sharp edges in the rivet joint to minimize exposed area
- Debonding study to analyze the effects of defects on delamination
 - Predicted electric potential in electrolyte and e-coat barrier
 - Recommended an e-coat for the sheet metal that would exhibit the lowest current and therefore the least decay



Top: Current density at the surface of a rivet and sheet metal Bottom: Electric potential in the e-coat and electrolyte in an electrocoated test sheet